

A Titration.

You are provided with a burette for sulphuric acid.

A 25 cm³ pipette for sodium hydroxide.

Phenolphthalein indicator.

You are to find the concentration of the sodium hydroxide solution in mol dm⁻³.

Pipette.

25.00 cm³ sodium hydroxide.

Burette.

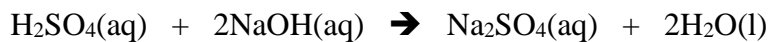
1.00 mol dm⁻³ sulphuric acid

Indicator.

Phenolphthalein colour change _____ to _____

Results.

	Trial	1	2	3
Initial vol. cm ³				
Final vol. cm ³				
Titre vol. cm ³				



Average of concordant titres =

1 mole of sulphuric acid reacts with _____ moles of sodium hydroxide

Thus moles of sulphuric acid added =

Use the equation to work out the moles of sodium hydroxide.

The concentration of sodium hydroxide solution is. _____