A Titration.

You are provided with a burette for sulphuric acid.
A 25 cm^3 pipette for sodium hydroxide.
Phenolphthalein indicator.

You are to find the concentration of the sodium hydroxide solution in mol dm^{-3}.

Pipette.
25.00 cm^3 sodium hydroxide.

Burette.
1.00 mol dm^{-3} sulphuric acid

Indicator.
Phenolphthalein colour change _____ to _____

Results.

	Trial	1	2	3
Initial vol. cm^3				
Final vol. cm^3				
Titre vol. cm^3				

$$H_2SO_4(aq) + 2NaOH(aq) \rightarrow Na_2SO_4(aq) + 2H_2O(l)$$

Average of concordant titres =

1 mole of sulphuric acid reacts with _____ moles of sodium hydroxide

Thus moles of sulphuric acid added =

Use the equation to work out the moles of sodium hydroxide.

The concentration of sodium hydroxide solution is._____